

11. Mixtures

the reference state

distributions of the molecules of a
binary mixture

average properties in a binary
mixture

MIXTURES:

The reference state —

Consider two states of a pure gas α and β

The difference in the Gibbs free energy of each state can be written as

$$G^\alpha - G^\beta = A^\alpha - A^\beta + P^\alpha V^\alpha - P^\beta V^\beta \quad (1)$$

From $\left(\frac{\partial A}{\partial V}\right)_{T,n} = -P$ we can derive (2)

$$A^\alpha - A^\beta = - \int_{\beta}^{\alpha} P dV \quad (3)$$

substituting this into the above eqn (1) and adding and subtracting $\int \frac{nRT}{V} dV$ from the right hand side gives:

$$G^\alpha - G^\beta = P^\alpha V^\alpha - P^\beta V^\beta - nRT \ln\left(\frac{V^\alpha}{V^\beta}\right) - \int_{\beta}^{\alpha} \left(P - \frac{nRT}{V}\right) dV \quad (4)$$

now replace by

$$P^\alpha = \frac{n}{V^\alpha} \quad P^\beta = \frac{n}{V^\beta} \quad z = \frac{PV}{nRT} \quad (5)$$

$$\text{and } dV = \frac{-n}{P^2} dp,$$

to get

$$\begin{aligned} \mu^\alpha - \mu^\beta &= RT(z^\alpha - z^\beta) + RT \ln\left(\frac{P^\alpha}{P^\beta}\right) \\ &+ \int_{P^\beta}^{P^\alpha} \frac{P - \frac{PRT}{z}}{P^2} dp \end{aligned} \quad (6)$$

Relating μ^β to a standard potential μ° , defined as the chemical potential of the perfect gas at a standard density $\rho^\circ = 1 \text{ \AA}^{-3}$ gives

$$\mu^\beta = \mu^\circ + RT \ln \frac{\lambda^\beta}{\rho^\circ} \quad (7)$$

If we let $\rho^\beta \rightarrow 0$ then $z^\beta = 1$ and we can replace the activity a^β by ρ^β and substitute the above eqn (7) into the preceding one (6) to give

$$\begin{aligned} \mu^\alpha - \mu^\circ &= RT(z^\alpha - 1) + RT \ln \left(\frac{P^\alpha}{\rho^\circ} \right) \\ &+ \int_0^{P^\alpha} \left(\frac{P - pRT}{p^2} \right) dp \end{aligned} \quad (8)$$

where $\mu^\alpha - \mu^\circ = \mu$ the chemical potential we will be referring to, so that we can dispense with the superscript α :

For a pure gas at pressure P and density ρ we can therefore calculate the chemical potential μ using this reference state, if we have the equation of state for it:

$$\mu = RT(z - 1) + RT \ln \left(\frac{P}{\rho^\circ} \right) + \int_0^P \left(\frac{P - pRT}{p^2} \right) dp \quad (9)$$

Now let us consider a binary mixture.

Use a simple, not so accurate virial equation of state with two virial coefficients only:

$$P = RT(\rho + \rho^2 B) \quad (10)$$

which we can easily extend to a mixture.

For a binary mixture, the virial eqn of state is

$$P = RT \left(\rho + \rho_1^2 B_{11} + 2\rho_1 \rho_2 B_{12} + \rho_2^2 B_{22} \right) \quad (11)$$

where B_{11} and B_{22} are the second virial coeffs of pure fluid 1 and pure fluid 2 respectively. B_{12} is a cross term, found experimentally for a particular binary mixture (that is, of fluid 1 and fluid 2). Assume all 3 virials are known as a function of temperature.

Subst (10) into (9) :

$$\mu = 2RTB\rho + RT \ln(P/p_0) \quad (12)$$

for a pure gas

The Helmholtz free energy of a fluid mixture if the virial eqn of state containing only 2 virial coeffs is used, is :

$$A = \sum_i n_i \mu_i^0 + RT \sum_i n_i \left[\ln \left(\frac{n_i RT}{P V} \right) - 1 \right] + \frac{RT}{V} \sum_i \sum_j n_i n_j B_{ij} \quad (13)$$

For a binary mixture,

$$\mu_1 = \left(\frac{\partial A}{\partial n_1} \right)_{n_2} \quad \text{and} \quad \mu_2 = \left(\frac{\partial A}{\partial n_2} \right)_{n_1} \quad (14)$$

$$\mu_1 - \mu_1^0 = 2RT(\rho_1 B_{11} + \rho_2 B_{12}) + RT \ln \left(\frac{P_1}{p_0} \right) \quad (15)$$

$$\mu_2 - \mu_2^0 = 2RT(\rho_2 B_{22} + \rho_1 B_{12}) + RT \ln \left(\frac{P_2}{p_0} \right) \quad (16)$$

where again we will be using $\mu_1 - \mu_1^0$ itself and $\mu_2 - \mu_2^0$ itself as the chemical potentials

in the simulations.

1. First step is to choose the equilibrium gas mixture (the bulk phase):

specify T, P_1, P_2

calculate P_{tot} using eq. (11)

calculate μ_1 using eq. (15)

μ_2 using eq. (16)

mole fraction in the bulk phase is

$$y_1 = P_1 / (P_1 + P_2)$$

2. Now ready to begin the simulation.

Parameters are: T, μ_1, μ_2, y_1

3. Markov chain is

choose create/annih

choose create/annih

displace a random molecule

a) choose to create or annihilate

choose 1 or 2 according to y_1



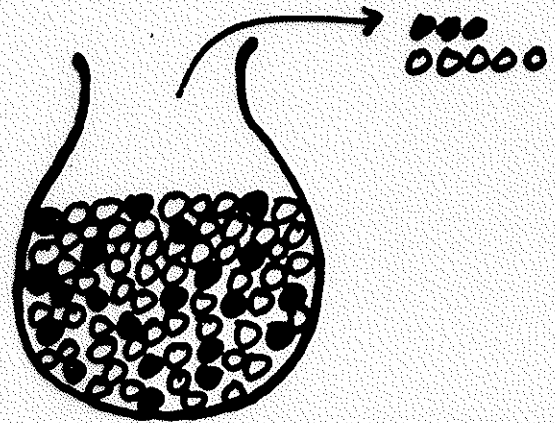
to ~~y_1~~

- What is the probability of drawing K balls n of them black out of an urn with M balls of which N_0 are black?

$$\binom{M}{K} \text{ or } C_M^K = \frac{M!}{K!(M-K)!}$$

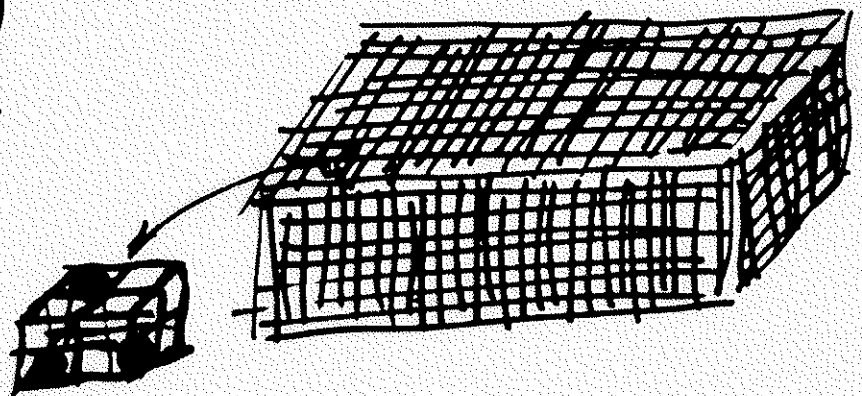
Answer:

$$P_n = \frac{\binom{N_0}{n} \binom{M-N_0}{K-n}}{\binom{M}{K}}$$



Exactly the same answer as the question
 What is the probability of finding a cell with n particles when there are a total of N_0 particles distributed among M sites, given that a site can only be occupied by one particle and that each cell has exactly K sites?

$$P_n = \frac{\binom{N_0}{n} \binom{M-N_0}{K-n}}{\binom{M}{K}}$$



Correspondence is
 each black ball represents an occupied site
 each white ball represents an empty site

Let both $M \rightarrow$ very large and $N_0 \rightarrow$ very large

What becomes of P_n ?

Let number of cells = Z
 let $\langle n \rangle$ be the average occupancy

$$M = KZ$$

$$N_0 = \langle n \rangle Z$$

$$P_n(\langle n \rangle) = \frac{\binom{N_0}{n} \binom{M-N_0}{K-n}}{\binom{M}{K}} = \frac{\binom{\langle n \rangle Z}{n} \binom{(K-\langle n \rangle)Z}{K-n}}{\binom{KZ}{K}}$$

$$= \frac{\langle n \rangle Z! \cdot [(K-\langle n \rangle)Z]!}{n! (\langle n \rangle Z - n)! \cdot (K-n)! \cdot [(K-2n)Z - K + n]!}$$

$$= \frac{(KZ)!}{K! (KZ - K)!}$$

Approximation:

$$\frac{N!}{(N-n)!} = \frac{N(N-1)(N-2)(N-3)\dots}{(N-n)(N-n-1)(N-n-2)\dots} = \underbrace{N(N-1)\dots(N-n+1)}_{n \text{ terms}}$$

$n \ll N$ so for very large N this is $\approx N^n$

$$\lim_{Z \rightarrow \infty} P_n(\langle n \rangle) = \frac{K!}{n! (K-n)!} \cdot \frac{(\langle n \rangle Z)^n [(K-\langle n \rangle)Z]^{K-n}}{(KZ)^K}$$

$$= \frac{K!}{n! (K-n)!} \cdot \frac{\langle n \rangle^n (K-\langle n \rangle)^{K-n}}{K^K} \cdot \left(\frac{Z^n Z^{K-n}}{Z^K} \right)$$

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$$= \frac{\langle n \rangle Z! \cdot [(K-\langle n \rangle) Z]!}{n! (\langle n \rangle Z - n)! \cdot (K-n)! \cdot [(K-\langle n \rangle) Z - K + n]!}$$

$$= \frac{(KZ)!}{K! (KZ-K)!}$$

Approximation:

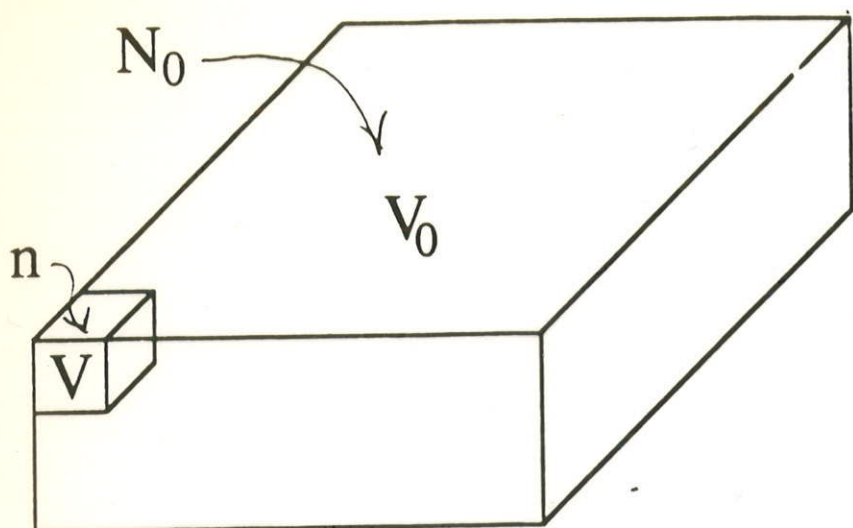
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The Hypergeometric Distribution



No. of zeolite cavities

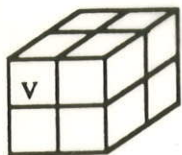
$$Z = \frac{V_0}{V} \sim 10^{20}$$

Total no. of sites

$$M = \frac{V_0}{v}$$

No. of sites per cavity

$$K = \frac{V}{v} \sim 8$$



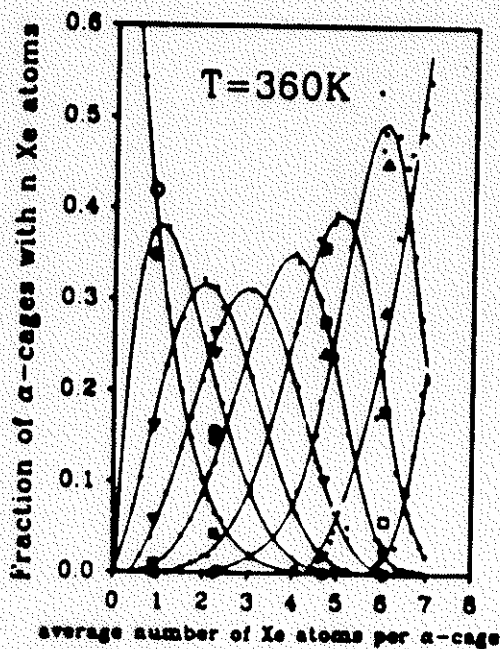
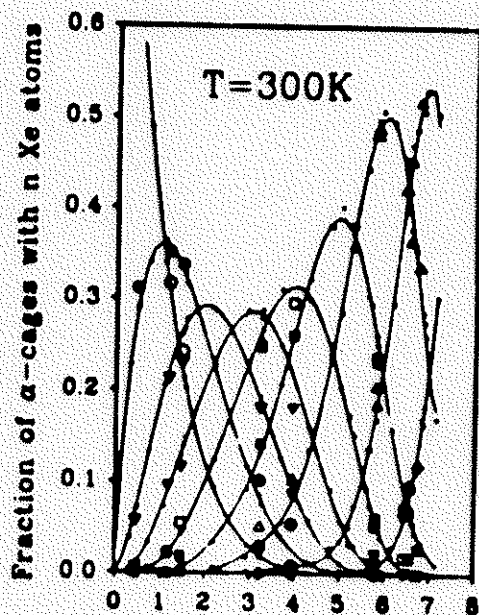
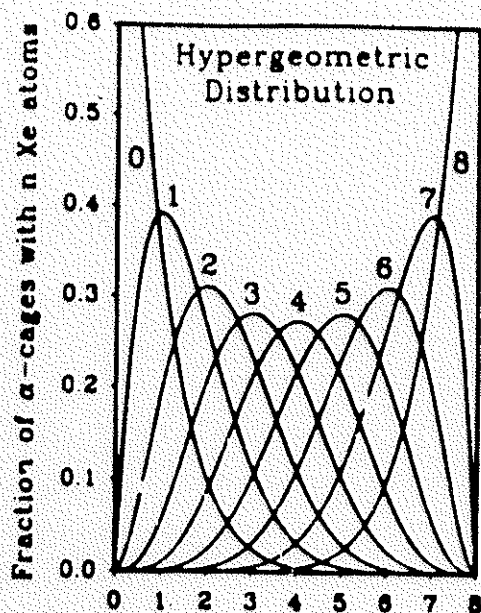
Average number atoms/cavity

$$\langle n \rangle = \frac{N_0}{(V_0/V)}$$

$$H_n(\langle n \rangle) = \frac{\frac{N_0!}{n!(N_0-n)!} \frac{(M-N_0)!}{(K-n)!(M-N_0-[K-n])!}}{\frac{M!}{K!(M-K)!}}$$

$$\lim_{Z \rightarrow \infty} H_n(\langle n \rangle) = \frac{\langle n \rangle^n (K - \langle n \rangle)^{(K-n)}}{K^K} \frac{K!}{n!(K-n)!}$$

How well does the hypergeometric distribution represent the actual distribution (EXPT) or distributions from "realistic" simulations?



The intrinsic probability of finding a blue particle in the mixture of blues and reds =

$$\frac{\langle i \rangle_{\text{blue}}}{\langle i \rangle_{\text{blue}} + \langle m \rangle_{\text{red}}}$$

The number of ways of arranging 4 blues and 3 reds is $\frac{7!}{4!3!}$

$$f(4, 3) = \frac{7!}{4!3!} H(7) \frac{\left(\frac{\langle i \rangle}{\langle i \rangle + \langle m \rangle}\right)^4 \left(\frac{\langle m \rangle}{\langle i \rangle + \langle m \rangle}\right)^3}{\frac{7!}{0!7!} \left(\frac{\langle i \rangle}{\langle i \rangle + \langle m \rangle}\right)^0 \left(\frac{\langle m \rangle}{\langle i \rangle + \langle m \rangle}\right)^7 + \frac{7!}{1!6!} \left(\frac{\langle i \rangle}{\langle i \rangle + \langle m \rangle}\right)^1 \left(\frac{\langle m \rangle}{\langle i \rangle + \langle m \rangle}\right)^6 + \dots + \frac{7!}{7!0!} \left(\frac{\langle i \rangle}{\langle i \rangle + \langle m \rangle}\right)^7 \left(\frac{\langle m \rangle}{\langle i \rangle + \langle m \rangle}\right)^0}$$

This simplifies to, in general, for $i + m = n$:

$$f(i, m) = H(n) \frac{\langle i \rangle^i \langle m \rangle^m}{i! m!} \frac{1}{\sum_{k=0}^n \frac{\langle i \rangle^k \langle m \rangle^{n-k}}{k! (n-k)!}}$$

The fraction of the cells that have exactly i blue particles and m red particles

$$\text{where } H(n) = \frac{\langle n \rangle^2 (8 - \langle n \rangle)^{8-n}}{8^8} \frac{8!}{n! (8-n)!}$$

= fraction of the cells having exactly n particles regardless of the number of red or blue among the n

DISTRIBUTIONS OF the molecules of a binary mixture. A SIMPLE MODEL

Question:

What is the probability of finding a cell with i blue particles and m red particles ($i+m=n$) when there are a total of N_0 particles of which I_0 are blue and M_0 are red?

Model: Assume that the particles are distributed among M sites, given that a site can only be occupied by one particle and that each cell has exactly K sites.

- For a binary mixture, assume that red and blue particles compete equally for the K sites in each cell.

Cast in the form of average occupancies

Z cells altogether

$M = KZ$ sites altogether (let $K=8$)

$$\langle i \rangle = \frac{I_0}{Z} \quad \langle m \rangle = \frac{M_0}{Z}$$

average number of blue particles per cell

average number of red particles per cell

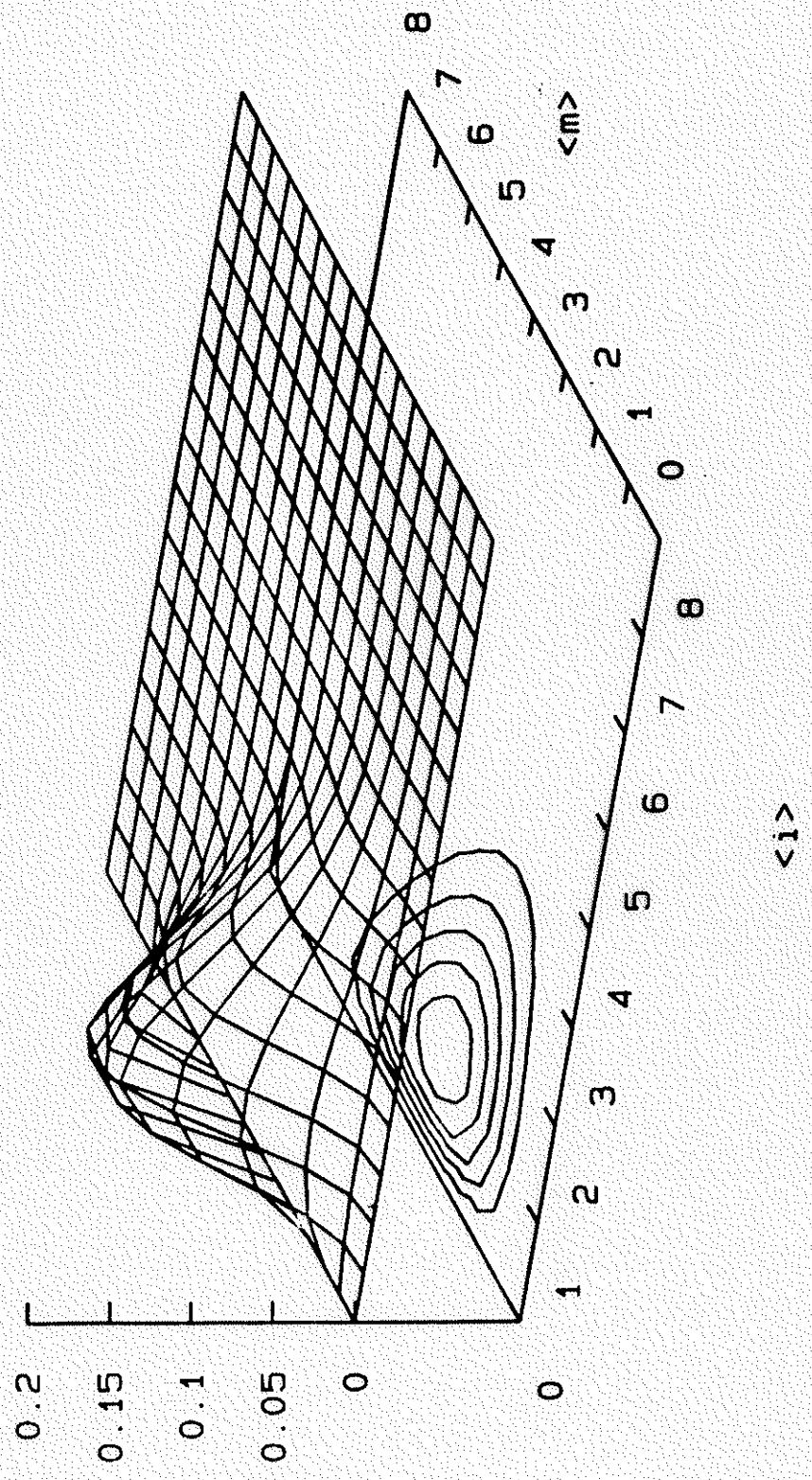
Distribution:

What is the fraction of cells that have exactly 4 blue and 3 red particles?

call this $f(4, 3)$
 ↑ ↑
 blue red

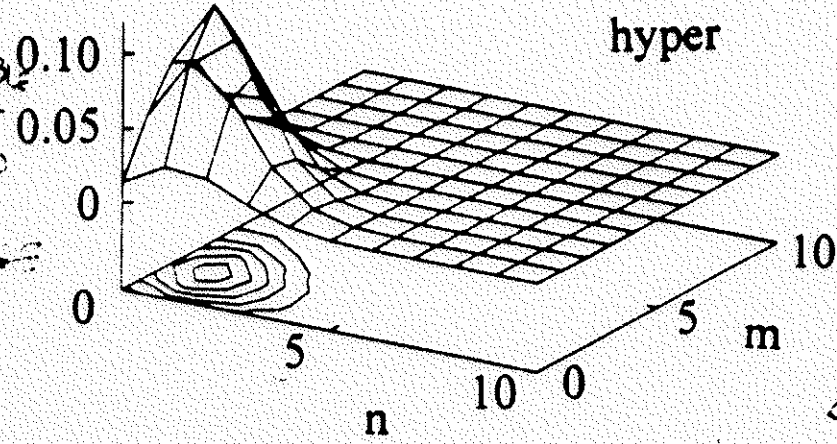
F (1, 3) IN ALPHA CAGES

1 3 —
1 3 —



COMPONENT ATOMS ARE DISTINGUISHABLE BUT EQUIVALENT IN COMPETITION FOR 2 LATTICE SITES PER CELL UNDER MUTUAL EXCLUSION

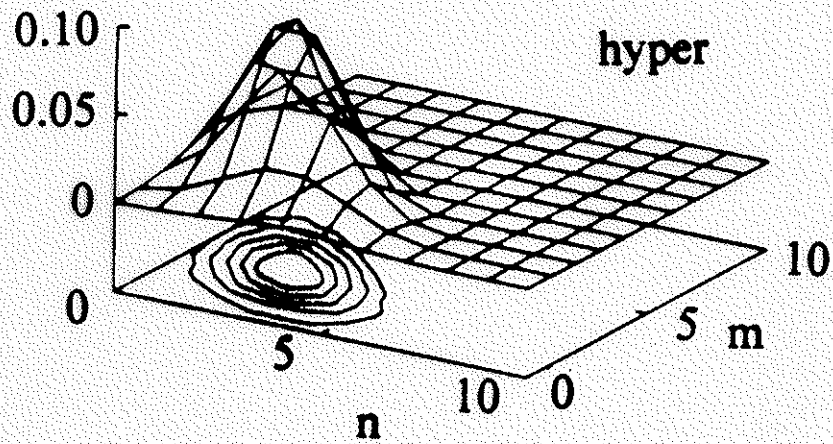
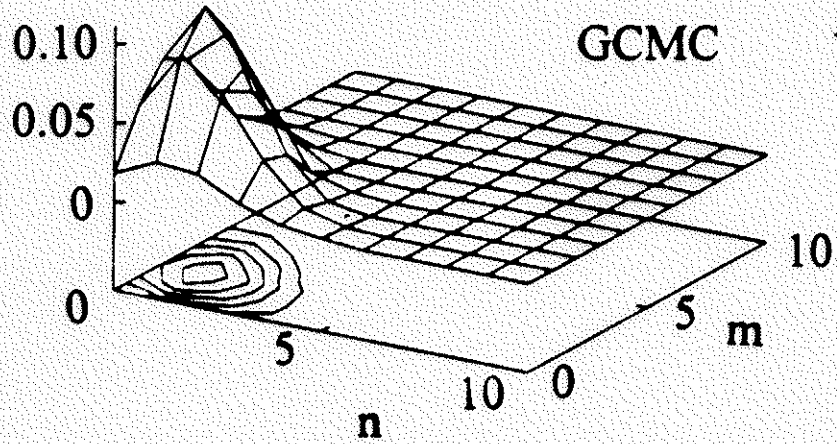
$$f(x_e, n, A r_m)$$



A SIMPLE MODEL for the DISTRIBUTION

$$\langle n \rangle_{x_e} = 2.232$$

$$\langle m \rangle_{Ar} = 2.083$$



$$\langle n \rangle_{x_e} = 2.429$$

$$\langle m \rangle_{Ar} = 3.140$$

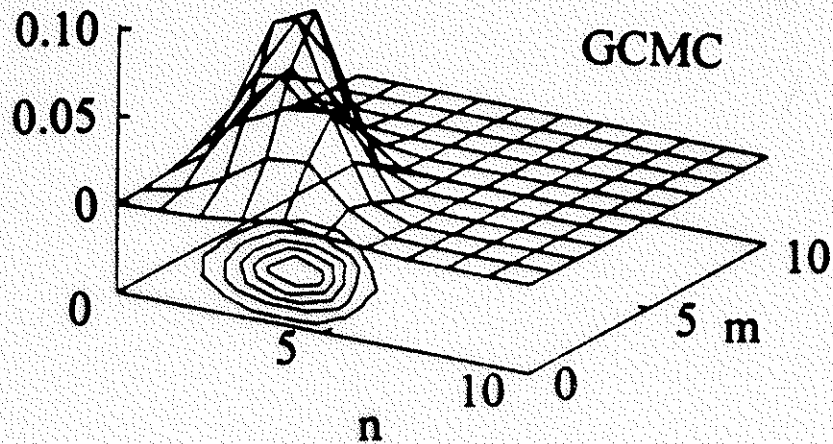
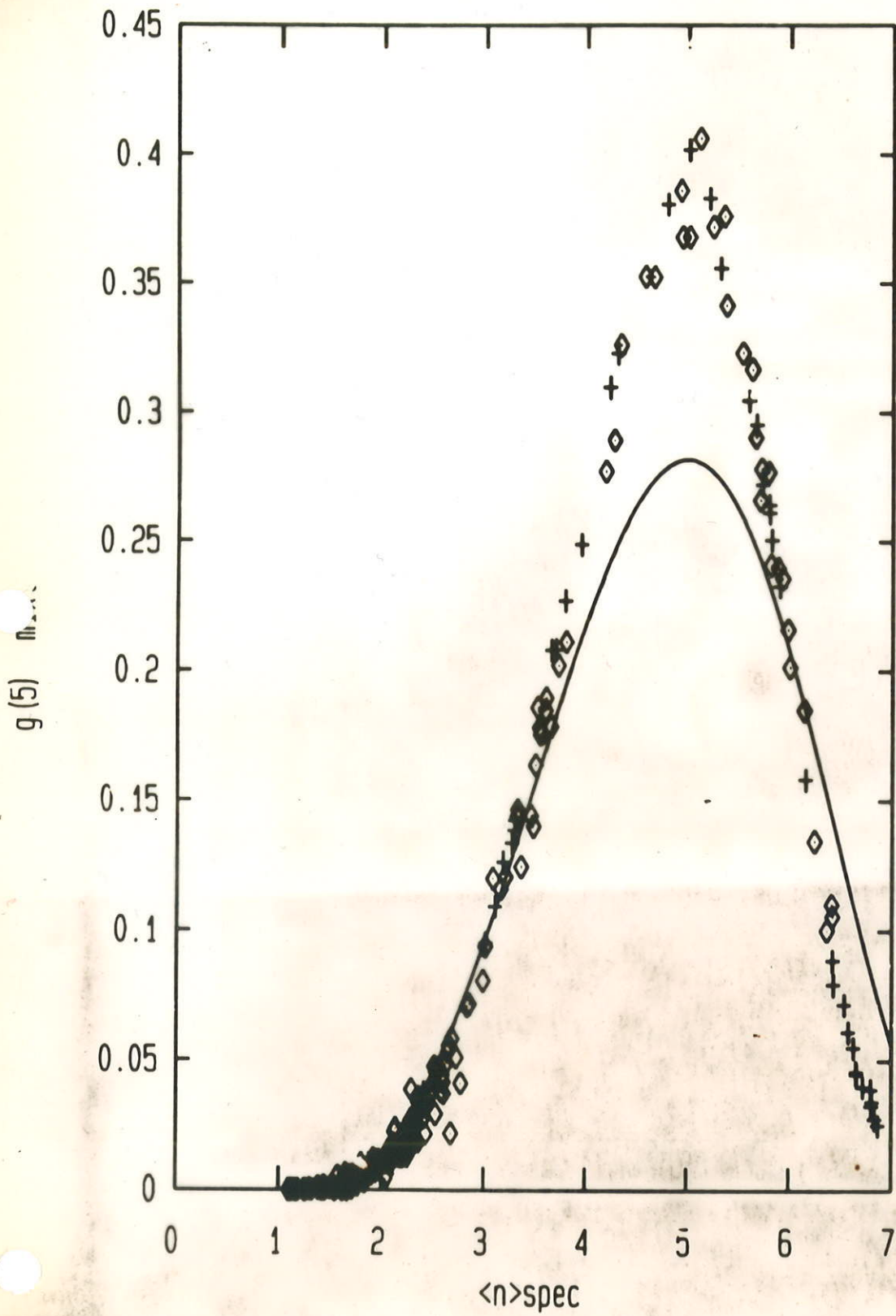


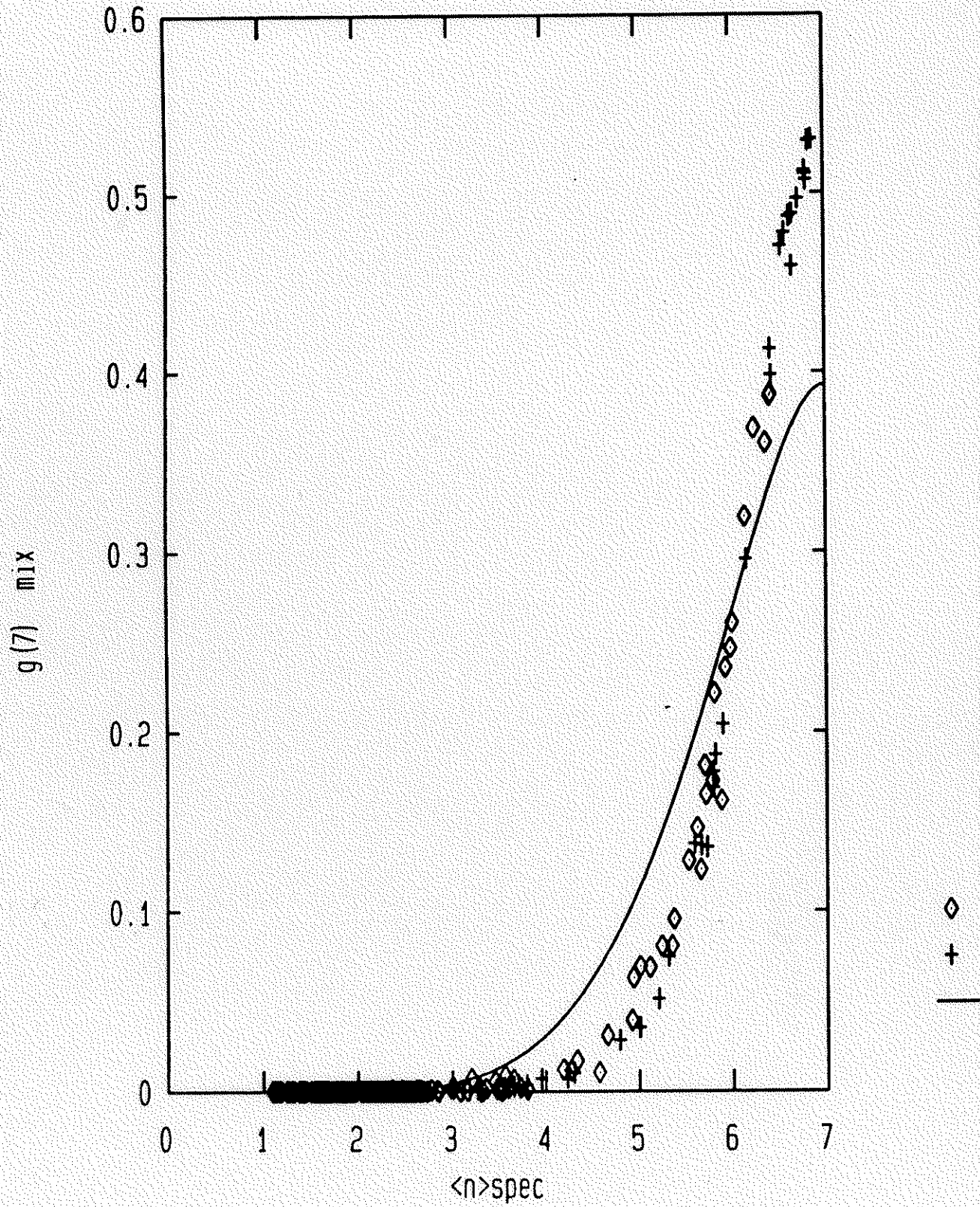
Table VI. Positive values of $[f(Xe_i, Ar_j) - f(Xe_j, Ar_i)]$ for a Xe-Ar mixture in zeolite NaA $\langle n \rangle_x$
 $\langle m \rangle_{Ar} = 3.65$, obtained from GCMC simulations. Shown are the number of Xe and Ar atom
 cluster and the values of $[f(Xe_i, Ar_j) - f(Xe_j, Ar_i)] > 0$.

1	1,0 none ^a	0,1 none																		
2	2,0 none		0,2 none																	
3	3,0 .00008	2,1 00005	1,2 0,3																	
4	4,0 .0006	3,1 00114		1,3 0,4																
5	5,0 .0040	4,1 0107	3,2 0067	2,3 0,5																
6	6,0 .00567	5,1 0225	4,2 0338		2,4 1,5										0,6					
7	7,0 .00117	6,1 00923	5,2 0230	4,3 0110											1,6				0,7	
8	8,0	7,1	6,2	5,3											2,6 0194	3,5 0142			1,7 0038	0,8 0001
9	9,0	8,1	7,2	6,3											3,6 0220	4,5 0170			2,7 0153	1,8 0022

Distribution of Xe in NaA



Distribution of Xe in NaA



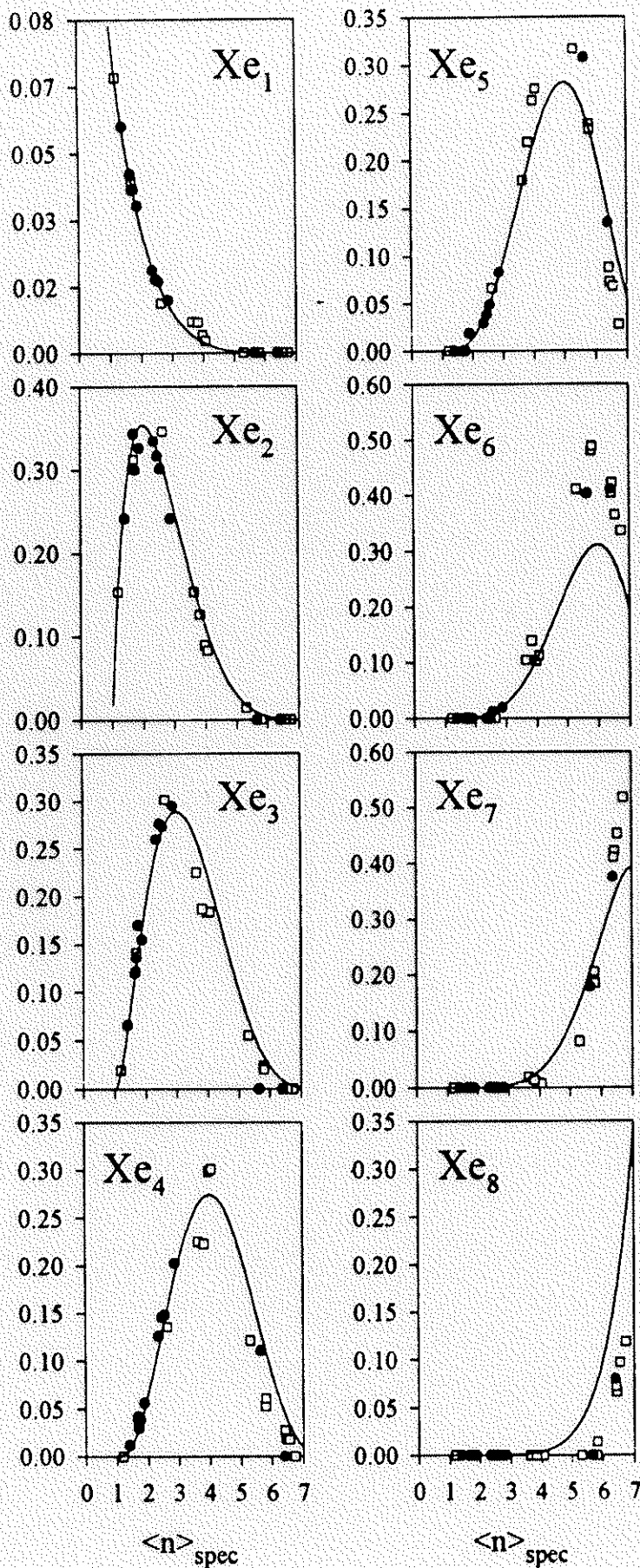


Fig. 14. Experimental equilibrium distribution of Xe atoms at 300 K among those cages occupied by Xe atoms. Shown are the fraction $g(n)$ of cages containing Xe_n in samples of zeolite NaA containing: (\square) pure Xe and (\bullet) a mixture of Xe and Ar. The fractions predicted by the hypergeometric distribution for 8 equivalent lattice sites is shown as the solid curve in each case.

We can use the fractions $f(i, m)$ to calculate averages such as: for a constant i

$$\langle \alpha(Xe_i; A_{\text{average}}) \rangle = \frac{\sum_{m=0} \sigma(Xe_i; A_m) i f(i, m)}{\sum_{m=0} i f(i, m)}$$

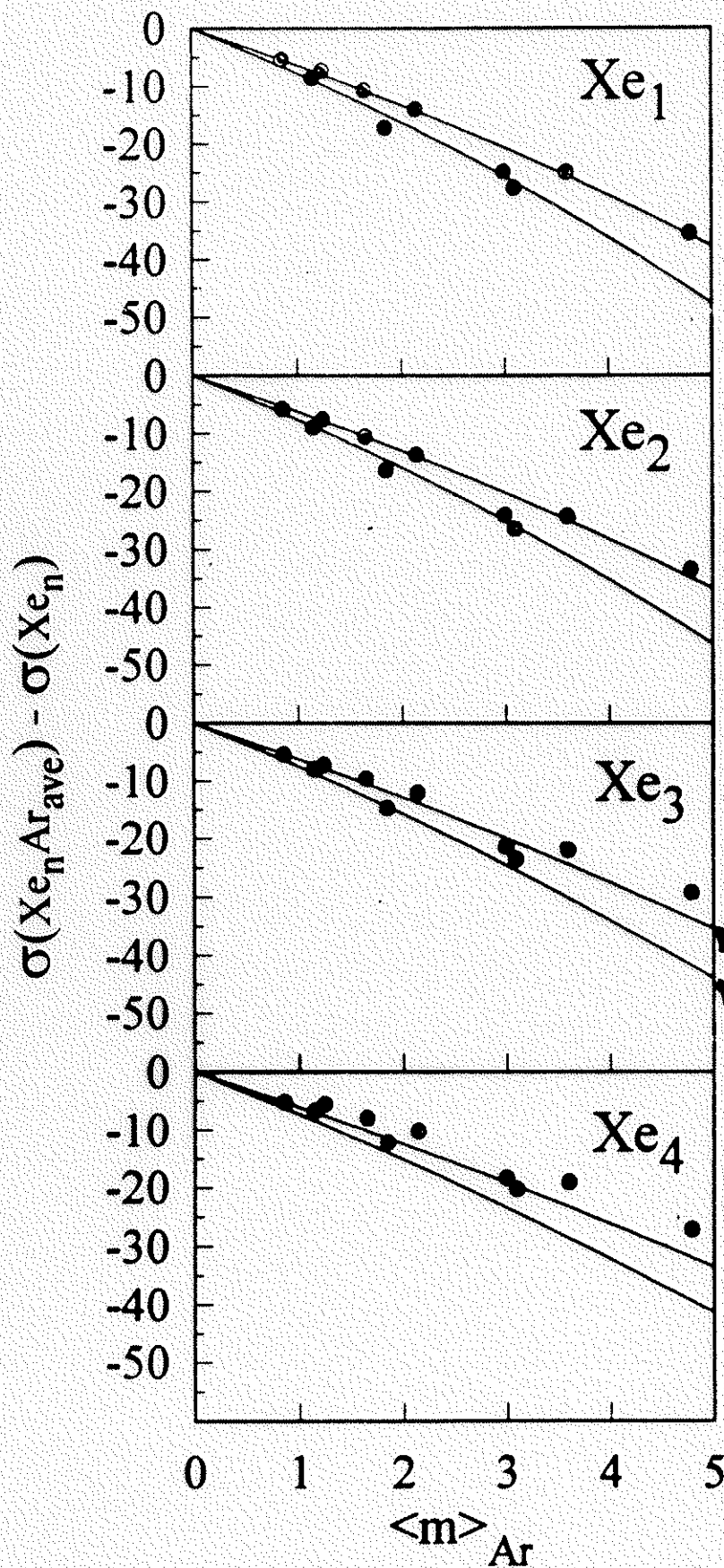
The fraction of cells containing i Xe atoms is $P_i = \sum_{m=0} f(i, m)$

Red and blue particles competing equally for the 8 sites in a cell means that

$$f(n, m) = f(m, n) \quad \text{for all } m \text{ and } n \text{ combinations}$$

$\begin{matrix} \uparrow & \uparrow & & \uparrow & \uparrow \\ \text{blue} & \text{red} & & \text{blue} & \text{red} \end{matrix}$

That is, the particles are equivalent.



The points are
EXPERIMENTAL
DATA

$$\langle n \rangle_{Xe} = 0.81 - 1.54$$

$$\langle n \rangle_{Xe} = 2.16 - 2.77$$

The curves are
based on
simple model
distribution
for $\langle n \rangle_{Xe} = 1.00$
and $\langle n \rangle_{Xe} = 2.40$

using
 $\sigma(Xe_n Ar_m)$
mixed cluster